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1 a) The relative atomic mass of an element is the weighted average mass of the isotopes of the element. (1) It is measured on a scale on which a carbon-12 atom has a mass of exactly 12 units. (1), b) The total mass of 100 atoms = (75 × 35) + (25 × 37) = 3550 (1) The average mass of 1 atom = 3550/100 = 35.5 (1)

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