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POWER LAW BASICS | 6 2 POWER LAW BASICS 2.1 ARRHENIUS EQUATION The main assumption behind the Arrhenius expression is that $k = A e^{-E_a/RT}$. This is an approximation, but it works quite well. The rate coefficient is the term that is a function of temperature but may also depend

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Chemical Equilibrium $dG = -SdT + VdP + \sum_{j=1}^n \mu_j dn_j$; μ_j : chemical potential for species j.
 $\mu_i(T, P) = \mu_i(T, P^\circ) + RT \ln \frac{P_i}{P^\circ}$, $\mu_j = 0$, $i = 1, \dots, n$
 $\mu_j = G_j + RT \ln a_j$, $a_j = f_j/f_j^\circ$
 $K_i = \prod_{j=1}^n Y_j^{\nu_j}$
 $\mu_i = -RT \ln K_i$ Standard state: pure species j at 1 atm and system temperature.

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Description. Chemical kinetics is the study of the rates of chemical reactions; if changes in conditions impact the speed of a reaction, we can better understand what caused the reaction. In this course, you will learn to understand complicated data sets and analysis techniques for measuring and understanding these rate changes.

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Grigoriy Yablonsky is an expert in the area of chemical kinetics and chemical engineering, particularly in catalytic technology of complete and selective oxidation, which is one of the main

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driving forces of sustainable development. His theory of complex steady-state and non-steady state catalytic reactions, is widely used by research teams in many countries of the world. Now, Grigoriy Yablonsky serves as an Associate Research Professor of Chemistry at Saint Louis University's Parks ...

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