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Electrons In Atoms

138 Chapter 5 • Electrons in Atoms

Although the speed of all electromagnetic  
waves in a vacuum is the same, waves can



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have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

Chapter 5: Electrons in Atoms

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## Electrons In Atoms

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Match. Gravity. Created by.

SmileyKylie0923. Key Concepts: Terms in this set (57) Dalton. The atom is a tiny, indestructible particle with no internal structure. Thomson. The atom is a sphere of positive electrical charge with electrons embedded in the sphere.

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Wave-Particle Duality/Electromagnetic

Spectrum/Relationship of

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Wavelength, Frequency and Speed of light  
5.2 Bohr's Model of the Atom/Quantum  
Mechanical Model of the Atom 5.3 Electron  
Arrangement & Valence Electrons.

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## Electrons In Atoms

the Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford ' s model fails to explain why objects change color when heated.

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Section 5.2 – Electron Arrangement in

Atoms The electron configuration of an  
atom is the arrangement of the electrons.

There are 3 rules that govern the electron  
configuration: Aufbau ' s principle, Pauli  
Exclusion principle, and Hund ' s rule.

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Chapter 5 “ Electrons in Atoms ”

Chemistry Charles Page High School

Stephen L. Cotton \* \* \* \* \* The

electromagnetic spectrum consists of radiation over a broad band of wavelengths. The visible light portion is very small. It is in the  $10^{-7}$ m wavelength range and  $10^{15}$  Hz (s<sup>-1</sup>) frequency range.



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## Electrons In Atoms

Unit 4: Electrons in Atoms (Chapter 5)

You'll Remember ...

Chapter 5 Electrons in Atoms. STUDY. PLAY. Quantum Mechanical Model. model of the atom we believe today that involves the probability of finding an electron in a certain position. What is the maximum number of f orbitals in any single energy

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level in an atom ? 7. Electrons in the same orbital.

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CHAPTER 5 What You ' ll Learn You will compare the wave and particle models of light. You will describe how the frequency of

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light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical

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138 Chapter 5 Electrons in Atoms Electron  
Configurations for Elements in Period Three

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Table 5-4 Figure 5-19. This sublevel diagram shows the order in which the orbitals are usually filled. The proper sequence for the first seven orbitals is 1s, 2s, 2p, 3s, 3p, 4s, and 3d. Chapter 5 Electrons in Atoms Flashcards | Quizlet

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Chapter 5: Electrons in Atoms Models of the Atom

- Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun.

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Chapter 5: Electrons in Atoms Models of the Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford ' s model fails to



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explain why objects change color when heated.

Chapter 5 Electrons In Atoms Workbook  
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CHAPTER 5 What You ' ll Learn You will compare the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and con-*trast* the Bohr and

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quantum mechanical models of the atom.  
You will express the arrangements of  
electrons in atoms through orbital

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How many electrons can each p orbital  
hold? Chapter 5: Electrons in Atoms

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Chapter 5 Electrons in Atoms 2. Light and Quantized Energy (5.1) <ul><li>The study

of light led to the development of the quantum mechanical model.

</li></ul><ul><li>Light is a kind of electromagnetic radiation (EM).

</li></ul><ul><li>All move at  $3.00 \times 10^8$

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m/s (c) Speed of light. </li></ul> 3.

## Answers

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