

2005 Ap Chem Frq Answers

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~~NaOH(aq). (i) Calculate the number of moles of NaOH(aq) added. molNaOH= 6.55 mL × 1 L 0.435 mol NaOH 1,000 mL 1 L × molNaOH= 2.85 × 10⁻³mol NaOH. One point is earned for the correct number of moles of NaOH. AP®CHEMISTRY 2005 SCORING GUIDELINES (Form B) Copyright © 2005 by College Board. All rights reserved.~~

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chemical equation $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons \text{HI}(\text{g})$, we get the reaction quotient expression: $Q_c = \frac{[\text{HI}]^2}{[\text{H}_2][\text{I}_2]}$ With $Q_c = 5.56$, $[\text{H}_2] = 0.030\text{M}$, and $[\text{I}_2] = 0.015\text{M}$, we get $[\text{HI}] = 0.050\text{M}$. b) We know that at equilibrium, $[\text{H}_2] = 0.020\text{M}$.

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